1st Semester Organic Chemistry
Exam 1

Name:

Recitation Instructor:

Instructions

1. Keep the exam closed until you are instructed to begin.

2. The exam consists of 9 questions. The first thing you should do is make sure that no pages are missing. If a page is missing, notify a proctor immediately.

3. You will have 1 hr and 15 minutes to complete the exam, at which time pencils must be put down. Budget your time wisely.

4. Questions are ranked easy (*), medium (**), and hard (***). You probably shouldn't spend too much time on a hard question if you still have easy or medium difficulty problems you haven't tried yet.

4. Make sure to show all of your work, and make it clear what your thought process was. Answers should fit in the space provided. If you need to use the back of the sheet of paper, you must make note of it in the space allotted for credit.

5. GOOD LUCK!

Breakdown

1 ____/10

2 ____/10

3 ____/10

4 ____/10

5 ____/10

6 ____/10

7 ____/20

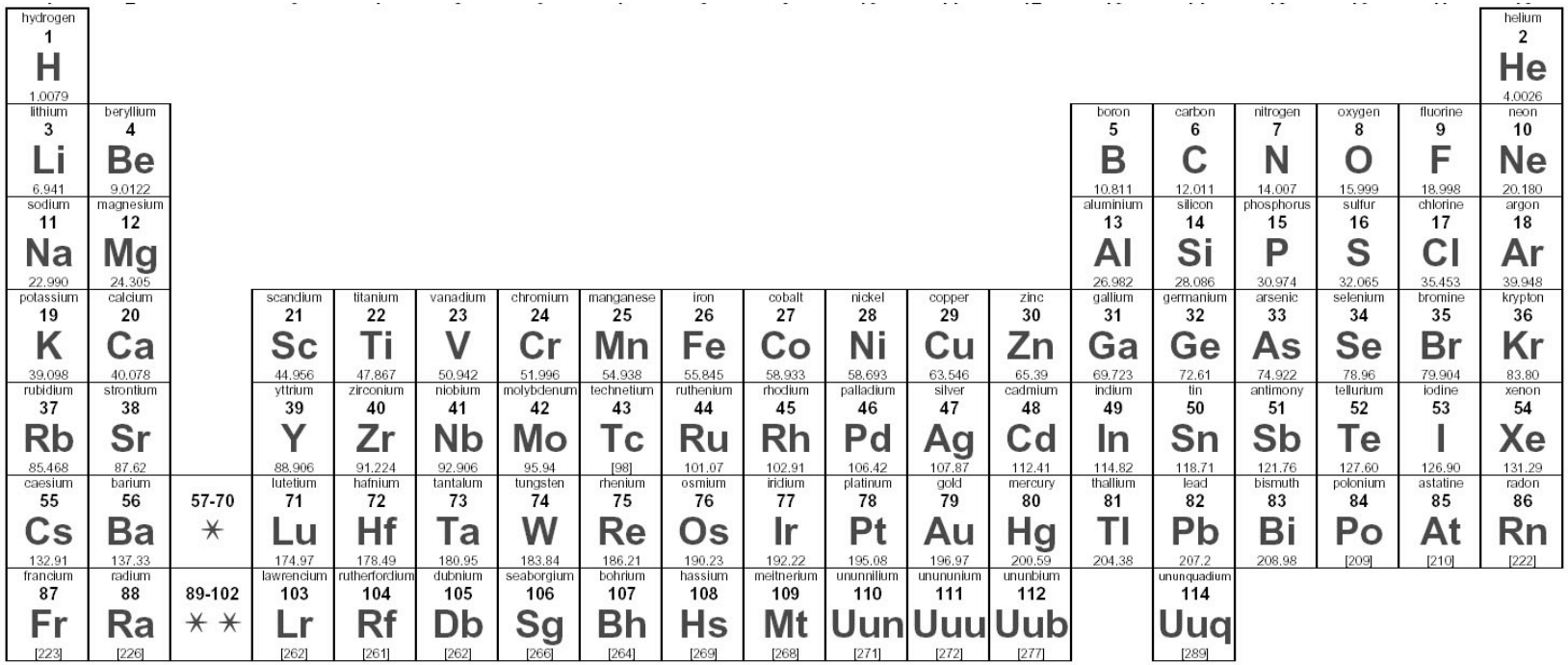
8 ____/10

9 ____/10

Tot. _____/100

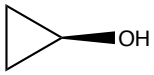
"Of course the elements are earth, water, fire and air. But what about chromium? Surely you can't ignore chromium."

From: sciencecartoonsplus.com


hydrogen 1 H 1.0079
helium 2 He 4.0026
lithium 3 Li 6.941
beryllium 4 Be 9.0122
boron 5 B 10.811
carbon 6 C 12.011
nitrogen 7 N 14.007
oxygen 8 O 15.999
fluorine 9 F 18.998
neon 10 Ne 20.180
sodium 11 Na 22.990
magnesium 12 Mg 24.305
aluminium 13 Al 26.982
silicon 14 Si 28.086
phosphorus 15 P 30.974
sulfur 16 S 32.065
chlorine 17 Cl 35.453
argon 18 Ar 39.948
potassium 19 K 39.098
calcium 20 Ca 40.078
scandium 21 Sc 44.956
titanium 22 Ti 47.867
vanadium 23 V 50.942
chromium 24 Cr 51.996
manganese 25 Mn 54.938
iron 26 Fe 55.845
cobalt 27 Co 58.933
nickel 28 Ni 58.693
copper 29 Cu 63.546
zinc 30 Zn 65.39
gallium 31 Ga 69.723
germanium 32 Ge 72.61
arsenic 33 As 74.922
selenium 34 Se 78.96
bromine 35 Br 79.904
krypton 36 Kr 83.80
rubidium 37 Rb 85.468
strontium 38 Sr 87.62
yttrium 39 Y 88.906
zirconium 40 Zr 91.224
niobium 41 Nb 92.906
molybdenum 42 Mo 95.94
technetium 43 Tc [98]
ruthenium 44 Ru 101.07
rhodium 45 Rh 102.91
palladium 46 Pd 106.42
silver 47 Ag 107.87
cadmium 48 Cd 112.41
indium 49 In 114.82
tin 50 Sn 118.71
antimony 51 Sb 121.76
tellurium 52 Te 127.60
iodine 53 I 126.90
xenon 54 Xe 131.29
caesium 55 Cs 132.91
barium 56 Ba 137.33
57-70 ✶
lutetium 71 Lu 174.97
hafnium 72 Hf 178.49
tantalum 73 Ta 180.95
tungsten 74 W 183.84
rhenium 75 Re 186.21
osmium 76 Os 190.23
iridium 77 Ir 192.22
platinum 78 Pt 195.08
gold 79 Au 196.97
mercury 80 Hg 200.59
thallium 81 Tl 204.38
lead 82 Pb 207.2
bismuth 83 Bi 208.98
polonium 84 Po [209]
astatine 85 At [210]
radon 86 Rn [222]
francium 87 Fr [223]
radium 88 Ra [226]
89-102 ✶ ✶
lawrencium 103 Lr [262]
rutherfordium 104 Rf [261]
dubnium 105 Db [262]
seaborgium 106 Sg [266]
bohrium 107 Bh [264]
hassium 108 Hs [269]
meitnerium 109 Mt [268]
ununnilium 110 Uun [271]
unununium 111 Uuu [272]
ununbium 112 Uub [277]
ununquadium 114 Uuq [289]
*Lanthanide series
lanthanum 57 La 138.91
cerium 58 Ce 140.12
praseodymium 59 Pr 140.91
neodymium 60 Nd 144.24
promethium 61 Pm [145]
samarium 62 Sm 150.36
europium 63 Eu 151.96
gadolinium 64 Gd 157.25
terbium 65 Tb 158.93
dysprosium 66 Dy 162.50
holmium 67 Ho 164.93
erbium 68 Er 167.26
thulium 69 Tm 168.93
ytterbium 70 Yb 173.04
**Actinide series
actinium 89 Ac [227]
thorium 90 Th 232.04
protactinium 91 Pa 231.04
uranium 92 U 238.03
neptunium 93 Np [237]
plutonium 94 Pu [244]
americium 95 Am [243]
curium 96 Cm [247]
berkelium 97 Bk [247]
californium 98 Cf [251]
einsteinium 99 Es [252]
fermium 100 Fm [257]
mendelevium 101 Md [258]
nobelium 102 No [259]

1. IUPAC (10 points, 5 points each)

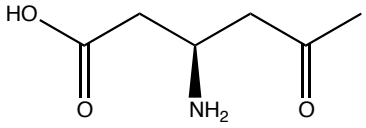
a. Give the name of the following molecule.* _____



b. Draw bicyclo[3.3.1]nonane using line-angle notation.*

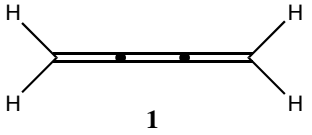
2. Functional Groups (10 points)

Circle and name all functional groups on the following molecule.*

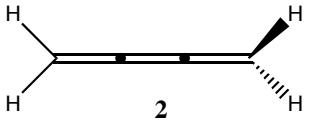


3. Hybridization (10 points)

Which of the following stereochemical representations of the extended allene are most accurate (**1** or **2**). Explain your answer, drawing in molecular orbitals (p orbitals) to help explain your answer.**



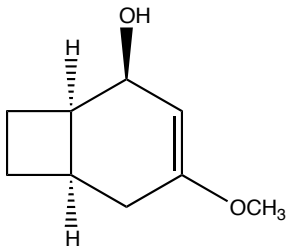
1



2

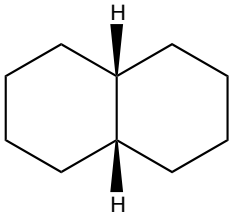
4. Stereochemistry Part 1 (10 points)

Circle the stereocenter(s) (ie 'chiral centers') on the following molecule and designate it (them) as R or S.**

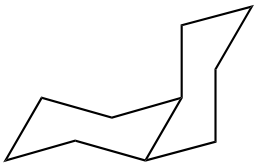


5. Stereochemistry Part 2 (10 points)

Cis dekalin drawn in its planar form has a plane of symmetry, suggesting that it is achiral. However, when drawn in the chair-chair form, the molecule does not have a super-imposable mirror image, suggesting that it is chiral. Would you classify *cis* dekalin chiral or achiral? Use structures where possible to support your answer. ***



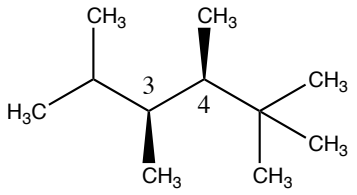
planar *cis* dekalin



chair-chair *cis* dekalin

6. Conformational Analysis Part (10 points)

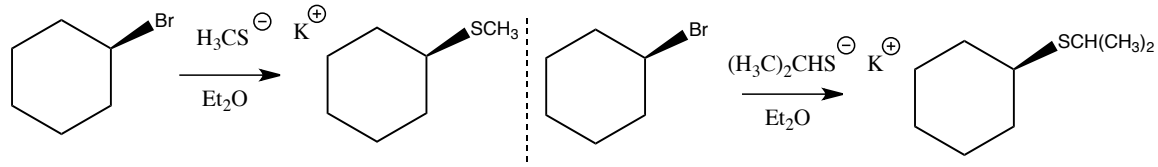
Draw all 3 staggered conformations of the following molecule in a Newman Projections down the 3-4 bond. Predict which would be the <u>lowest</u> in energy and explain your answer. **



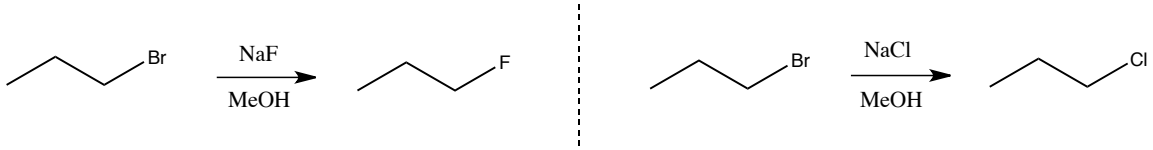
7. Substitution Reactions (20 points, 5 points each)

For each of the following substitution reactions, predict which reaction would be <u>faster</u> and explain your answer.

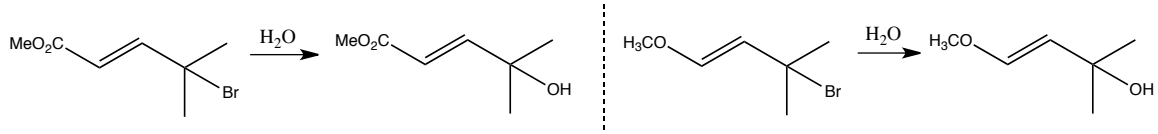
7a. SN2*



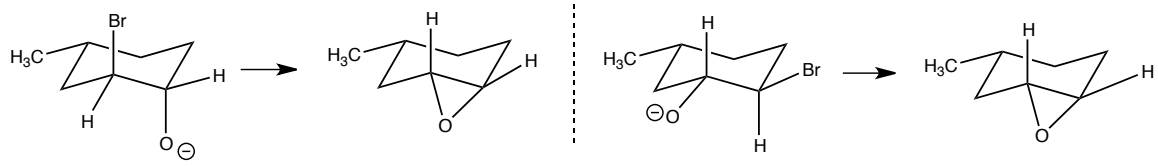
7b. SN2**



7c. SN1**

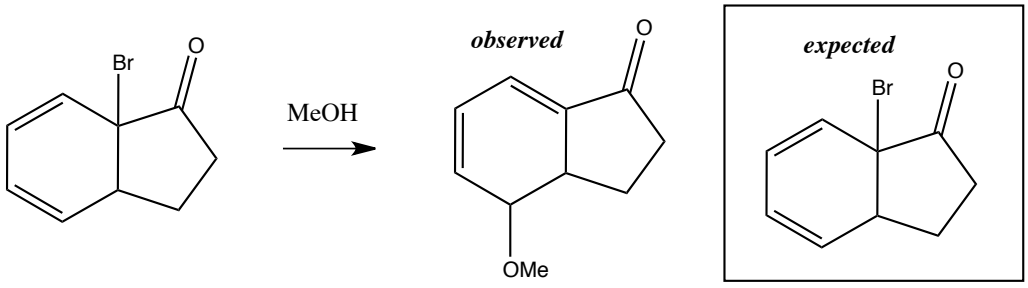


7d. SN2**



8. Mechanism # 1: Substitution (10 points)

Researchers were attempting to carry out an SN1 reaction to generate the expected product shown. However, instead a different product was observed. Propose a mechanism for the reaction that leads to the <u>observed</u> product.**



9. Mechanism # 2: Resonance (10 points)

Show two resonance forms that illustrate how the positive charge generated through the protonation of the xanthene molecule can be shared with the two other oxygens. Use reaction arrows to show this interconversion.***